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- 3. What is meant by the term Reaction Rate?
- 4. What is the collision theory?
- 5. What is the activation energy?6. What is a catalyst and why is it different
- from a reactant in an equation?7. What are the FOUR major factors that
- affect reaction rate?8. How does each of the factors above affect the rate of reaction?
- Draw an exothermic reaction graph shown with and without a catalyst?
- 10. Why would a mixture of gases react faster when the volume they occupy is decreased?
- 11. Why would iron filings rust faster than an iron nail?
- 12. How would the increasing pressure of reactive components of a gaseous mixture affect the rate at which the components react with one another?
- 13. How would you change temperature of a reaction if you wanted to increase the rate of reaction? Explain how this effects the reaction using the collision theory.
- 14. Why would the rate of reaction decrease as the reaction produces more products?
- 15. If you put 100g of NaOH in cube form and 200g of NaOH in powdered form which will react with HCl at a faster rate and list the reasons why?
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